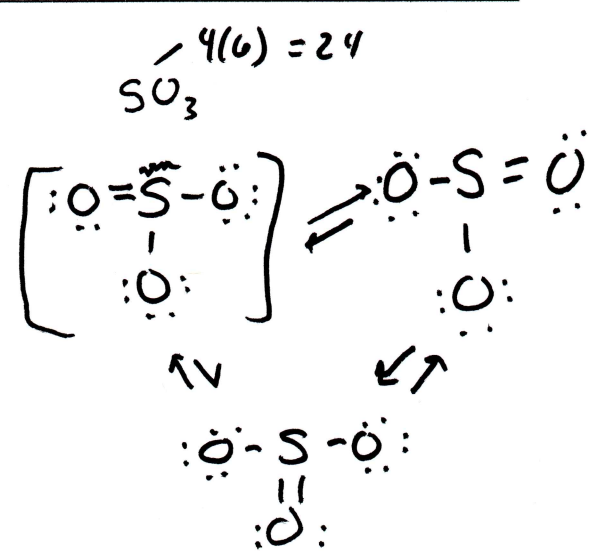


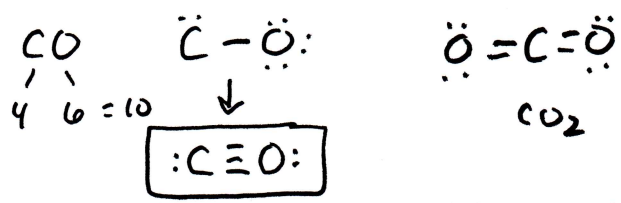
Topic Reminder Q6
Covalent Bonding and Bond properties

Complete the following nomenclature questions

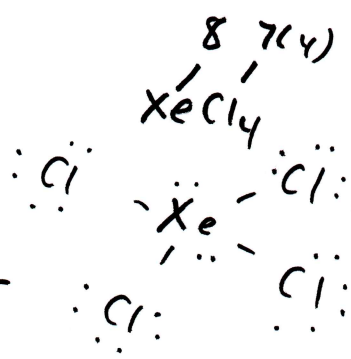
- CO = Carbon Monoxide
- SO₃ = Sulfur trioxide
- Xenon tetrachloride = XeCl₄
- Carbon dioxide = CO₂



5. Draw a Lewis structure of each substance above.



6. ~~CO~~ is polar. Draw a partial charges on the molecule.



7. What is the name of each molecular structure?

- CO Linear XeCl₄ - Square planer
 SO₃ - Trigonal planer CO₂ - Linear

8. Are any of the substances above violating the octet rule?

yes XeCl₄ Xe has 12 e⁻

9. Which of the substances above has lots of London Dispersion Forces due to lots of electrons.

XeCl₄ or SO₃

10. Below you will find a water molecule. Add another water molecule demonstrating hydrogen bonding. Add a ... symbolizing the hydrogen bond.

